

CORRECTION.

The Relative Solubility of the Silver Halides and Silver Sulphocyanate.—In Table I of this paper (THIS JOURNAL, 30, p. 72) the ratio of the solubility of silver chloride to silver sulphocyanate appears in an inverted form. The mean ratio calculated from the table should read

$$\frac{S_{\text{AgCNS}}}{S_{\text{AgCl}}} = 0.0748.$$

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NEW BOOKS.

A Text-Book of Electro-Chemistry. By MAX LE BLANC. Translated from the Fourth Enlarged German Edition by WILLIS R. WHITNEY and JOHN W. BROWN. New York: The Macmillan Company. Price, \$2.60 net.

Since the first edition of Le Blanc's treatise appeared in 1895 it has been accepted the world over as the standard text-book of electrochemistry. In the succeeding editions the author has aimed to keep pace with the rapidly growing science. It is doubtful whether the resulting growth of the volume to more than double the original size has increased the value of the book as a text-book, but it has provided a remarkably handy and comprehensive compendium of electrochemical knowledge.

The style remains the style of a text-book, and the subject will seem easier to the reader than it is in reality. Weak points in the theory are slurred over. The recent critique of Jahn, the question as to the correctness of the ionization values calculated from the conductivity, the enormous deviation of strong electrolytes from the mass law, receive but scant attention. It has been the misfortune of the ionic theory that its advocates have seldom been satisfied with pointing out its unquestionable triumphs, but have claimed for it a perfection which it has not yet attained.

The author has succeeded to an extraordinary degree in bringing his work up to date. Important investigations which appeared even up to within a few months of the date of publication of the book are mentioned and frequently their results are incorporated in the text.

Occasional misstatements occur, as the one on page 181 that "the 'relations' between the solution pressures of various metals are independent of the nature of the solvent, and, moreover, always possess the same value." As a rule, however, the statements are accurate and reliable. This, unfortunately, is not true of the last chapter, in which the so-called decomposition potentials are discussed. The greater part of this chapter is devoted to an attempted explanation of phenomena which have been shown to be as purely subjective as the N-rays of Blondlot.

The English edition is rather more than a translation. Explanatory and supplementary paragraphs have been added, many new illustra-